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Thermochemistry MULTIPLE CHOICE.  
Choose the one alternative that best completes the statement or answers the question.  $4\text{Al (s)} + 3\text{O}_2\text{ (g)} \rightarrow 2\text{Al}_2\text{O}_3\text{ (s)}$   
A) exothermic, released B) exothermic, absorbed.

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Question 1 The following equation shows the reaction that occurs when nitroglycerine explodes:  $4 \text{C}_3\text{H}_5\text{O}_9\text{N}_3 \rightarrow 12 \text{CO}_2 + 6 \text{N}_2 + \text{O}_2 + 10 \text{H}_2\text{O} + 1725 \text{ kcal}$  This reaction is:

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heat needed to raise the temperature of  
1 g of water by 1 C 2. \_\_\_\_ SI unit of  
energy 3. \_\_\_\_ quantity of heat needed to

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change the temperature of 1 g of a ...

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dioxide with oxygen.  $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$  ( $\Delta H^\circ_f \text{SO}_2(\text{g}) = -296.8$  kJ/mol;  $\Delta H^\circ_f \text{SO}_3(\text{g}) = -395.7$  kJ/mol)  
a. -98.9 kJ c. 197.8 kJ b. -197.8 kJ d. Not enough information is given. 5 ID: A Ch 17 Thermochemistry Practice Test Answer Section MATCHING 1. ANS: OBJ: 2. ANS: OBJ: 3.

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set (14) heat capacity. the amount of  
heat required to change the temp. of an  
object by exactly 1 degree C. law of  
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162 CHAPTER 6: THERMOCHEMISTRY To  
convert the answer to joules, we write:  
 $101.3 \text{ J} \cdot 0.18 \text{ L atm} \cdot 1 \text{ L atm} = - \cdot \times = \cdot \text{ w}$   
 $-18 \text{ J} \cdot 6.17$  An expansion implies an

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increase in volume, therefore  $w$  must be  $-325 \text{ J}$  (see the defining equation for pressure-volume work.) If the system absorbs heat,  $q$  must be  $+127 \text{ J}$ . The change in energy (internal

## **CHAPTER 6 THERMOCHEMISTRY**

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