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Note: Not all of the reactions will occur.  
For those that do not, write no reaction.

1. A piece of copper is dropped into a  
container of water. No reaction  
2. Liquid bromine is added to a container of  
sodium iodide crystals.  $\text{Br (l)} + 2\text{NaI(s)}$   
 $2\text{NaBr(s)} + \text{I}_2\text{(s)}$   
3. An aluminum strip is  
immersed in a solution of silver nitrate.

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Practice-Pg.54 • Weak acids and bases  
do not break apart and are in net ionic  
equation. • Synthesis and  
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Handbook SnCO tin(II) carbonate sodium  
hydrogen carbonate NaHCO 1. 2. 3. 4. 5.  
6. 8. 9. 10. 11. vanadium(V) oxide H<sub>2</sub>O  
dihydrogen monoxide (NH) CO  
ammonium oxalate 4 224 polonium(VI)  
thiocyanate Po(SCN)<sub>6</sub> tetraphosphorus  
decaoxide P<sub>4</sub>O<sub>10</sub> 12. 13. 15. 16. 17.  
18. 19. manganese(VII) oxide copper(II)  
dihydrogen phosphate

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Inc. Correlating questions (ii) from:  
Rodney Schmitz (2007) Moore County  
Schools, NC Reaction Prediction - 5 For  
each of the following eight reactions, in

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part (i) write a BALANCED equation and in part (ii) answer the question about the reaction.

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